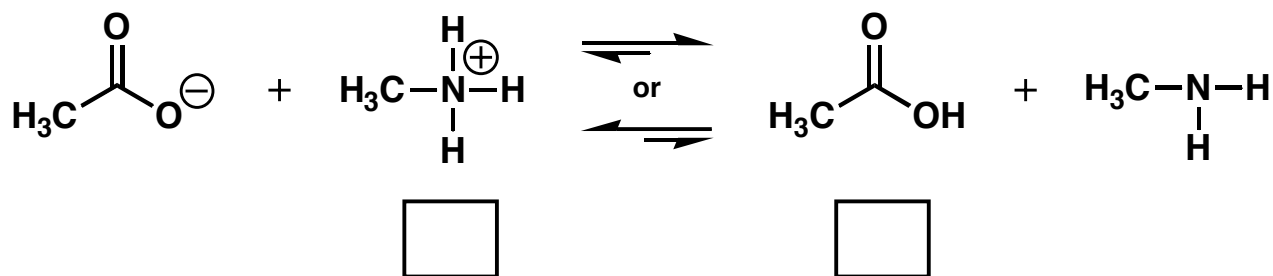
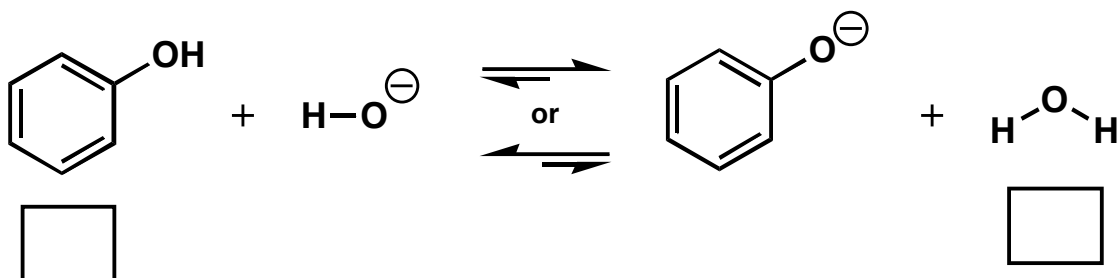


1. (1 pt) What is the pKa of a *sp* hybridized carbon atom?
- a) 25 b) 35 c) 45 d) 50
2. (2 pts) **a)** Circle the letter of the molecule with the lowest barrier to rotation (or ring-flip). **b)** Box the letter of the molecule with the highest barrier to rotation (or ring-flip).
- a) ethane
b) propane
c) butane
d) cyclohexane
3. (2 pts) Rank the following substituents in order of priority (1 = highest priority).
- _____ -CH=CH₂
_____ -CN
_____ -CH₂NH₂
_____ -CH₂Br
4. (1 pt) If a chiral molecule has an absolute configuration of R, which direction does it rotate the plane of polarized light?
- a) clockwise (dextrorotatory)
b) counterclockwise (levorotatory)
c) it doesn't rotate the plane of polarized light
d) can't be determined from the information given

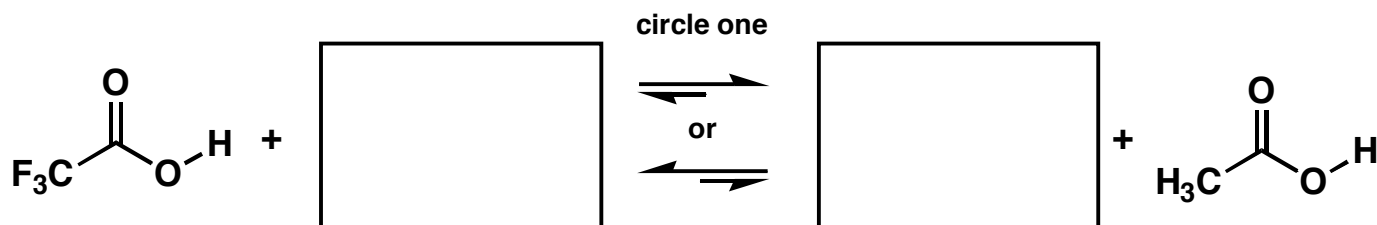
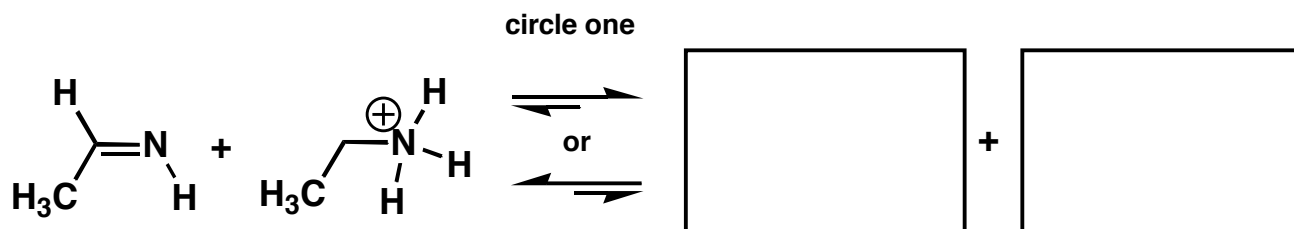
Initials

Points

5. (6 pts) a) Enter the pK_a value for each acid in the boxes below. b) Indicate whether the reactants or products will be favored at equilibrium by circling the appropriate set of equilibrium arrows (a longer arrow is drawn toward the species favored at equilibrium).



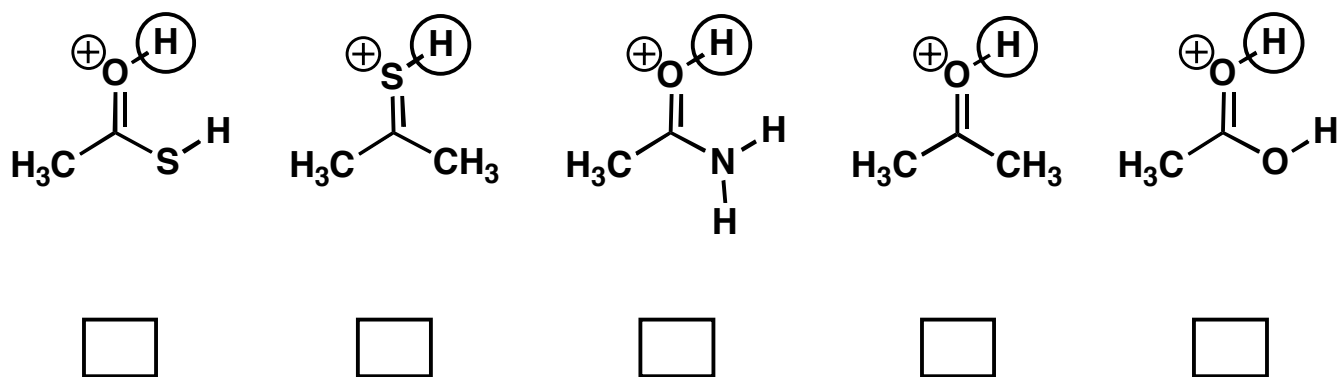
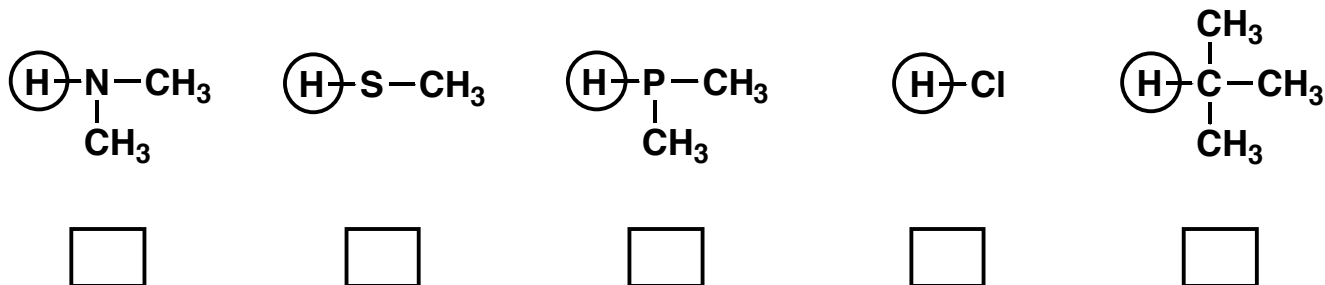
6. (10 pts) a) Provide structures in the boxes to complete the following acid-base (proton-transfer) reactions. b) For each set of reactants, draw in all lone pairs of electrons and show the electron movement by using curved arrows. c) Indicate whether the reactants or products will be favored at equilibrium by circling the appropriate set of equilibrium arrows.



Initials

Points

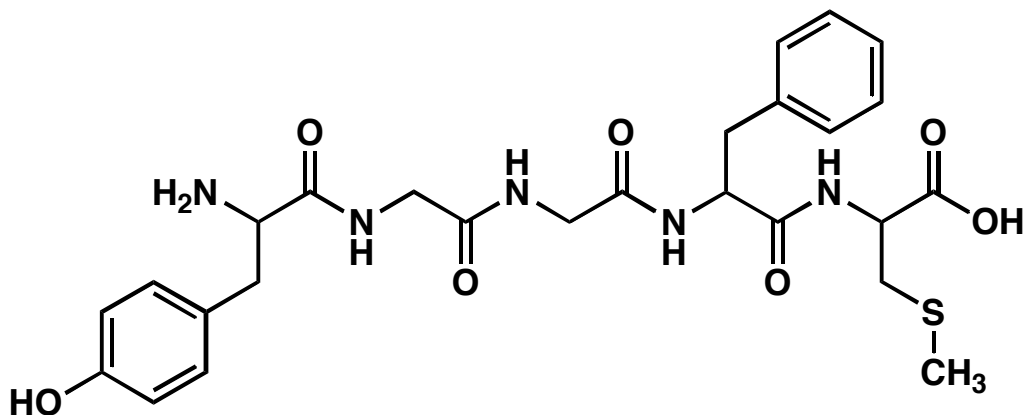
7. (10 pts) Rank the following sets of molecules in order of acidity (1 = most acidic).



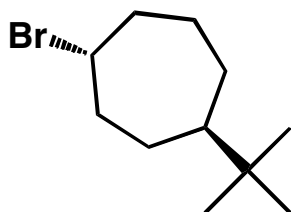
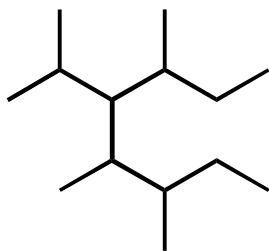
Initials

Points

8. (6 pts) Met-enkephalin, an endorphin, serves as a natural pain reliever that changes or removes the perception of nerve signals. Label all of the functional groups present in Met-enkephalin.



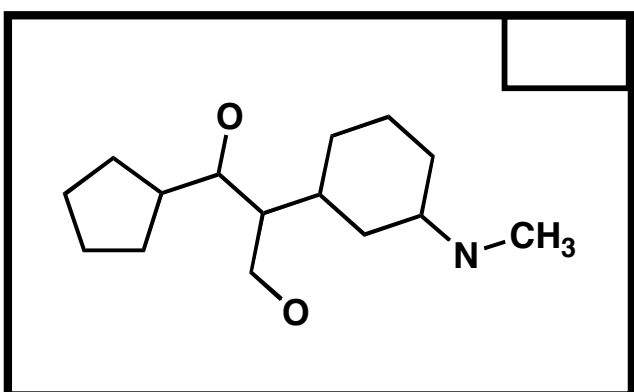
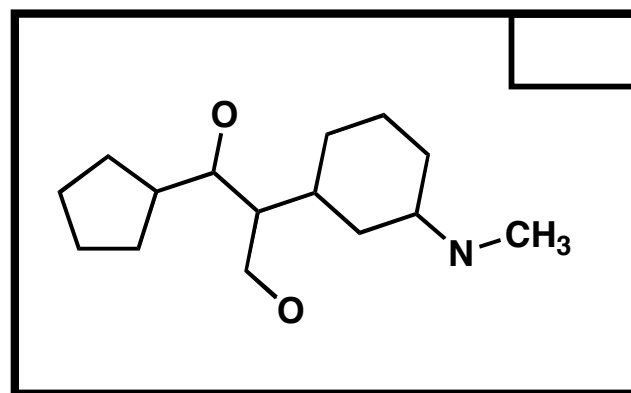
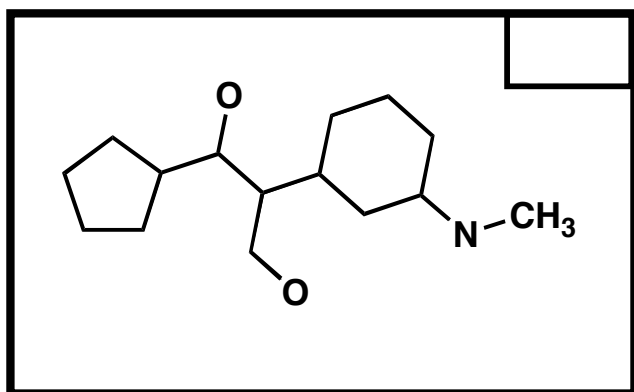
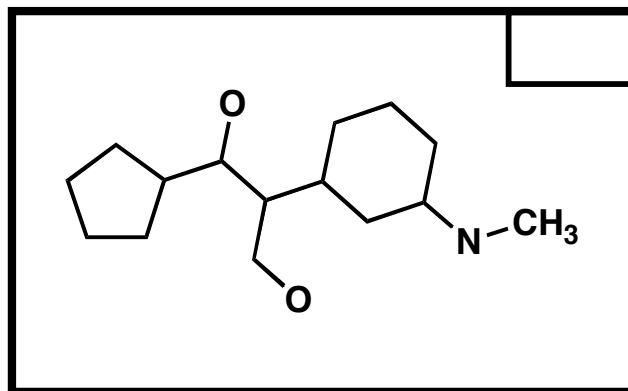
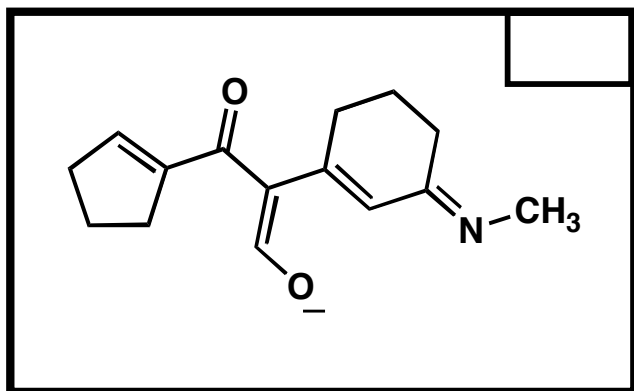
9. (4 pts) Name the following molecules.



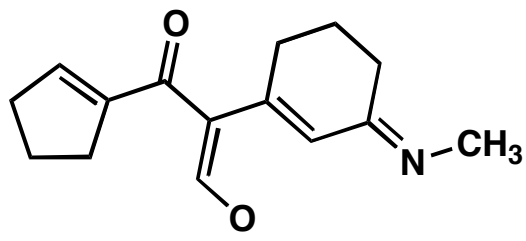
Initials

Points

10. (16 pts) **a)** Draw the 4 major resonance contributors for the molecule shown in the first box. Partially completed structures are provided as a time-saver. **Do not generate any additional charges.** **b)** Draw in all lone pairs of electrons and use arrows to show the movement of electrons within the structure. **c)** Place a checkmark in the small boxes of the three structures that contribute the most to the resonance hybrid. **d)** Circle all the nucleophilic atoms in the structure at the bottom of the page.



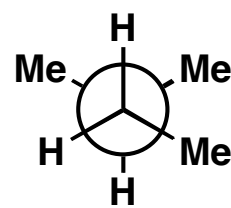
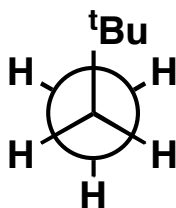
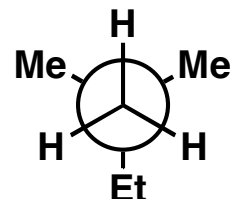
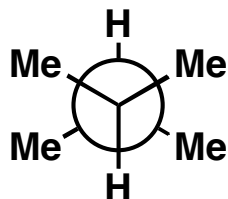
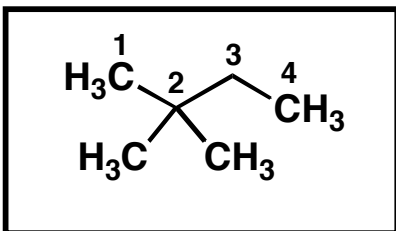
Circle all of the nucleophilic atoms in the molecule below.



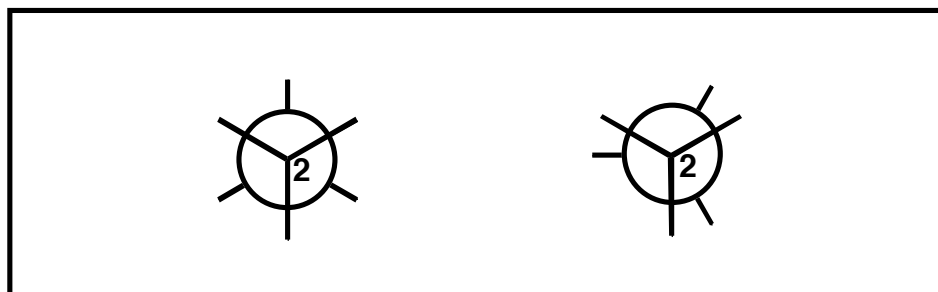
Initials

Points

11. (12 pts) a) Circle the structures that represent a conformation of 2,2-dimethylbutane sighting along **any** C-C bond.



b) Complete each of the Newman projections below to show the most stable and least stable conformations of 2,2-dimethylbutane, sighting along the **C₂-C₃** bond.



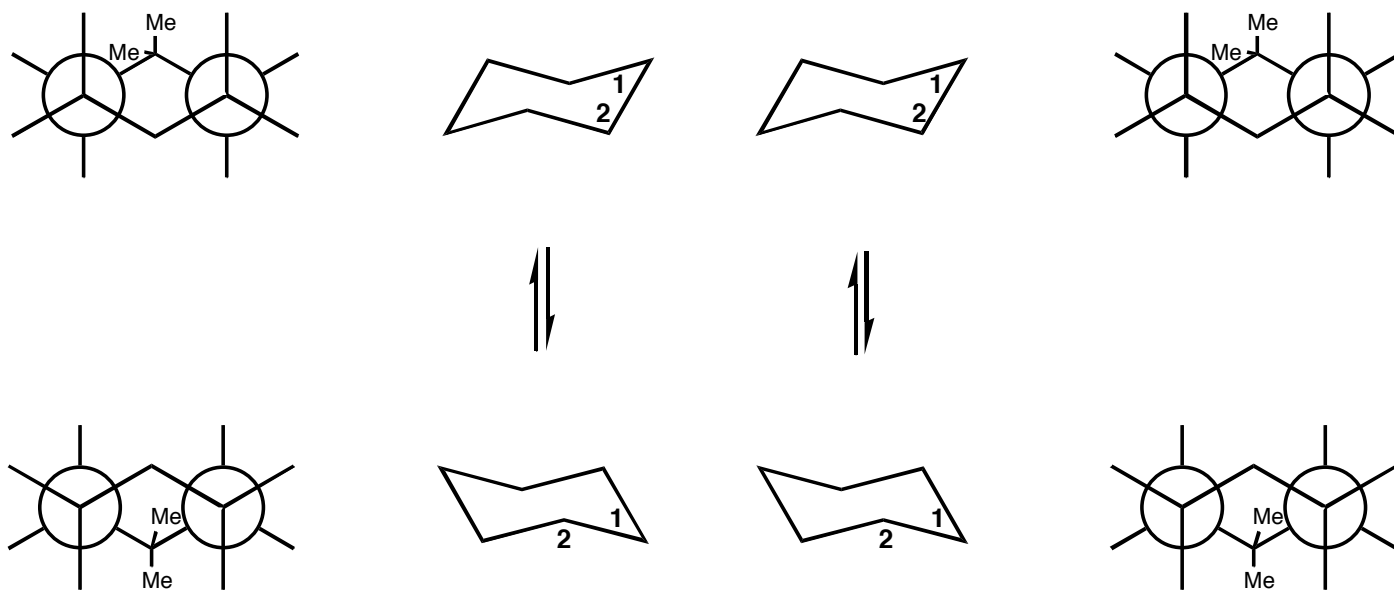
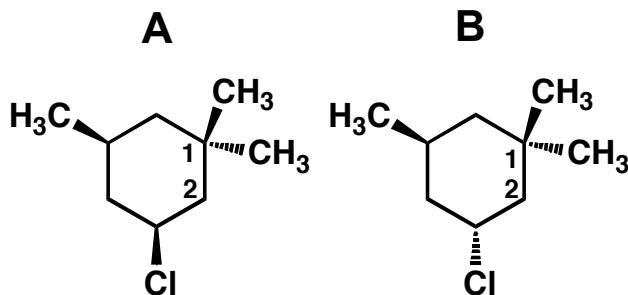
c) Use the above Newman projections to calculate the barrier to rotation of 2,2-methylbutane sighting along the **C₂-C₃** bond.

Barrier to Rotation = _____

Initials

Points

12. (12 pts) **a)** Draw in the substituents on the ring flipped conformers of each molecule (**A** and **B**) to predict which molecule is lower in energy. **b)** Show the Newman projection for each ring-flipped conformer, sighting along the C₅-C₆ and C₃-C₂ bonds.



c) Circle the chair conformation that is lower in energy for each molecule.

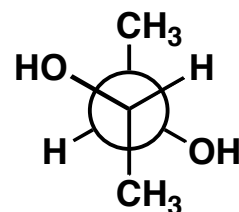
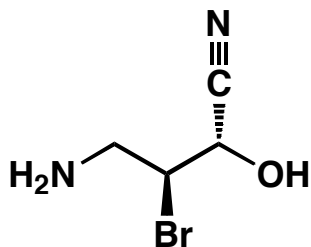
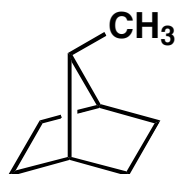
d) Which structure is lower in energy (circle one)? **A** or **B**

e) Briefly explain your choice:

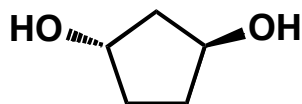
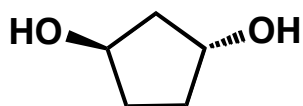
Initials

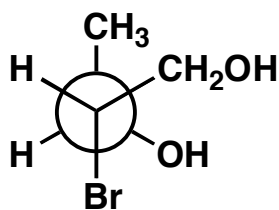
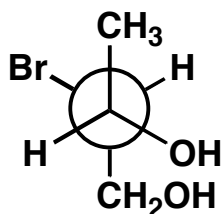
Points

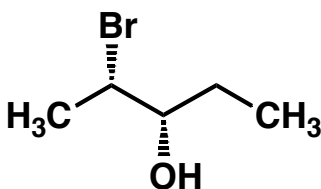
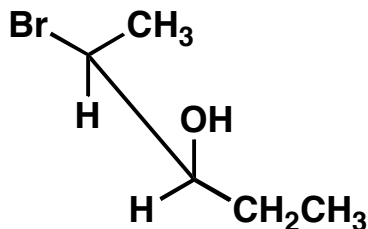
13. (9 pts) **a)** Label each molecule as chiral (**C**), achiral (**A**), or achiral/meso (**M**).
b) Designate each chirality center as **R** or **S**.



14. (9 pts) Indicate the relationships between the two molecules as enantiomers (**E**), diastereomers (**D**), or same molecule (**S**).







Initials

Points